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an atom, a formula unit, or a molecule. representative particle. the mass in grams of a mole of any pure substance. molar mass. 6.02×10^{23} representative particle of a substance. mole (mol) the molar mass of molecular oxygen is: 32g. at STP, 1 mol each of hydrogen, oxygen, nitrogen and fluorine (diatomic):

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Chemistry Chapter 10 Chemical Quantities. STUDY. PLAY. Mole. (mol) of a substance is 6.02×10^{23} representative particles of that substance and is the SI unit for measuring the amount of a substance. Avogadro's Number. the number of representative particles in a mole, 6.02×10^{23} . Representative Particle.

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the molar mass of molecular oxygen is. 32. at STP, 1 mol each of hydrogen, oxygen, nitrogen, and fluorine. have different densities. a large weather balloon filled with helium has a volume of 7.00×10^2 L at STP which expression should be used to find the mass. 7.0×10^2 L (1 mol/22.4L) (4g/mol)

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Chapter 10 Chemical Quantities91. SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289)

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

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