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oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2.

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Read Online Chapter 12 Assessment Stoichiometry Answer Key Read Online Chapter 12 Assessment Stoichiometry Answer Key Answer: 4.93×10^{-5} L or 49.3 μL In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para - nitrophenol. When the limiting reactant is

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