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Chapter 3. law of conservation of mass. law of definite proportions. law of multiple proportions. atom. States that matter can be neither destroyed nor created during.... The fact that a chemical compounds contains the same elements.... if two or more different compounds are composed of the same tw....

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Atoms. • Atoms are building blocks of all matter. • According to modern atomic theory, an atom is the smallest particle of an element which takes part in chemical reaction. • Atoms are very small and which can't be seen even through very powerful microscope. • Atomic radius is measured in nanometres.

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provided. 3 points each, 9 points total. 1. The element boron, B, has an atomic mass of 10.81 amu ...

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Chapter 3 Chemical Foundations: Elements, Atoms, and Ions
Chapter 11 Modern Atomic Theory 1 GPS: SC3 Students will use the modern atomic theory to explain the characteristics of 2 atoms. 0 a. Discriminate between the relative size, charge, and position of protons, 1 neutrons, and electrons in the atom. 2 b. Use the orbital configuration of neutral atoms to explain its effect ...

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CHAPTER 3 REVIEW

Atoms: The Building Blocks of Matter

SECTION 2 SHORT

ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

3 Atoms: The Building Blocks of Matter

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Mrs. J's Chemistry Page - Lecture Notes

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Chapter 3. Summary. Victor is now 17 years old and ready to become a student at the University of

Ingolstadt in Ingolstadt, Germany (near Munich), but an outbreak of scarlet fever at home delays his departure. His mother and "cousin" both fight the disease; Caroline Beaufort Frankenstein dies, and Elizabeth recovers.

Chapter 3

Taken from the book Modern Chemistry by Holt, Rinehart, and Winston on Chapters 4 and 5, which deals with electrons and the periodic table. Includes the chapter vocabulary and a few other useful things.

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