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Sulfur - 2, Oxygen - 8
3. $2\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$ 4.
Calcium - 4, Carbon -
4, Oxygen - 12 5. $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$ 6.
 $2\text{NaClO}_3 \rightarrow 2\text{NaCl} + 3\text{O}_2$
7. $8\text{Ag}_2\text{S} + 16\text{Ag} + \text{S}_8$
8. $\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow \text{CuS} + 2\text{HCl}$ 9. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ 10. $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$ 11.

Equation Challenge Answer Key - Science Spot

Answer Key Balancing
Equations

ChallengeBalancing
Equations Challenge

Part A: Parts & Pieces

(1) Circle each

subscript in each

chemical formula. (2)

Draw a square around
each coefficient. (3)

Answer the questions
related to each

chemical formula. Part

B: Label the chemical

equation using

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 2. Hydrogen - 4, Sulfur - 2, Oxygen - 8
 3. $2\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$
 4.

Equation Challenge Answer Key

T. Trimpe 2006 Answer Key Balancing

Equations Challenge

Balancing Equations

Challenge Part A: Parts & Pieces (1) Circle

each subscript in each chemical formula. (2)

Draw a square around each coefficient. (3)

Answer the questions related to each

chemical formula.

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complete combustion
of each challenge part
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character in the
element and lowercase
for the second
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