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relationship: as
temperature increases,
volume increases.

Boyle's law is an
inverse relationship: as
pressure increases,
volume decreases."

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A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant. Charles' Law Problems: $V_1T_2 = V_2T_1$. 1 atm = 760.0 mm Hg = 101.3 kPa. Calculate the decrease in temperature when 6.00 L at 20.0 °C is compressed to 4.00 L.

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Boyles And Charles

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