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 The Origin of the
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multiple choice

definition of self-

ionization $[H_3O^+]$ $[OH^-]$

$= 1 \times 10^{-14}$ definition

and formula for

calculating $pH = -$

$\log[H_3O^+]$ $pH > 7$ is a

base; $pH < 7$ is an acid;

$pH = 7$ is neutral pH

range is normally 0 to

14 calculate pH from

$[H_3O^+]$ calculate

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choice definition of

self-ionization $[H_3O^+]$

$[\text{OH}^-] = 1 \times 10^{-14}$
definition and formula
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Chapter 15 & 16 - Acid,

Base & Titration Terms.

Binary Acid. Oxyacid.

Arrhenius Acid.

Arrhenius Base. an acid

that does not contain

oxygen, such as

hydrofluoric acid. an

acid that is a

compound of

hydrogen, oxygen, and

a third el.... a

substance that

increases the

concentration of

HYDRONIUM ions....

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