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Figure 11.2 Three common chemical

reactions are shown. When methane gas burns, it combines with oxygen to form carbon dioxide and water. Iron turns to red-brown rust (iron(III) oxide) in the presence of oxygen in the air. Hydrogen peroxide decomposes to water and oxygen when used as an antiseptic.

11.1 Describing Chemical Reactions

11

11.4: Combination Reactions; 11.5: Decomposition Reactions; 11.6: Combustion Reactions; 11.7: Single Replacement Reactions; 11.8: Activity Series; 11.9: Double Replacement Reactions A double-replacement reaction is a reaction in which the positive and negative ions of two ionic compounds exchange places to form two new compounds.

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Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure

copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$ D. $\text{Fe (s)} + \text{Cl}_2\text{(g)} \rightarrow \dots$

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